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SECTION 2 SHORT

ANSWER Answer

the following

questions in the

space provided.

1. The following

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are statements about the dissolving process. Explain each one at the molecular level.

a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas in the

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A homogeneous mixture that contains particles in a dispersed phase that do not settle out is a solution. The solution process

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with water as
the solvent.
hydration.

Liquid solutes
and solvents
that are not
soluble in each
other are.

immiscible. A
solute molecule
that is
surrounded by
solvent
molecules is.

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review. STUDY.
PLAY. solution.
homogeneous
mixture of two
or more
substances,

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consisting of
ions or
molecules.
solute. gas or
solid dissolved
in a liquid or
solution
component in
small amount.
solvent. liquid
that dissolves a
gas or solid or
the solution
component in

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ANSWER Answer

the following

questions in the

space provided.

1. Match the

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type of mixture

on the left to

its

representative

particle

diameter on the

right. c

solutions (a)

larger than 1000

nm a suspensions

(b) 1 nm to 1000

nm b colloids

(c) smaller than

1 nm 2. Identify

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the solvent in
each of the
following
examples:

Answers

~~mc06se cFMs r i-
vi~~

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- 1) 56 2) = 4, =
10, = 16, = 22,
= 28 3) = 4, =
8, = , = 64, =
4) $a_n = 2n - 3$

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5) \bullet $4n$ 6) =

5, = 3, = 1, =

-1 7) $3 + 4 + 5$

+ ... + $(n + 2)$

8) 9) 15 10) $d =$

-5; = 4, = -1, =

-6, = -11 11) =

$94 - 10n; = 14$

12) 220 13) =

73, $d = -5, = -$

5 14) 34980 15)

25

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~~, = 10, = 16, =
22, = 28 + 1) =
2.~~

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SHORT ANSWER

Answer the
following
questions in the
space provided.

1. Describe the
errors made by

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the following
students in
making molar
solutions. 3.
James needs a
0.600 M solution
of KCl. He
measures out
0.600 g of KCl
and adds 1 L of
water to the
solid. James
made several
errors.

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SECTION 3 SHORT

ANSWER Answer

the following

questions in the

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12 Review provided.

1. Describe the errors made by the following

students in making molar solutions. a.

James needs a 0.600 M solution of KCl. He measures out 0.600 g of KCl and adds 1 L of water to the

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b. Mary
needs a 0.02 M
solution of
NaNO₃. She
calculates that
she needs 2.00

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learners to show
what they know.

They answer
questions

comparing

solutions,

suspensions, and
colloids.

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INSTRUCTIONS:

Write your
answers on lined
or blank sheets
of paper. You do
not need to write
the

question, ONLY

WRITE THE

PROBLEM NUMBER

and ANSWERS/SOLU

TIONS. • For

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involve
calculations,
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you must show
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your work to get
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full credit.

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of Mixtures
Solutions are
homogeneous
mixtures of two
or more
substances in a
single phase.
Soluble
describes a

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substance capable of being
dissolved.
Solvent is the
dissolving
medium in a
solution.

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Solutions are

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Solvent is the dissolving medium in a solution. Solute is the substance that is dissolved in a solution.

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