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Notes (answers)

Chapter 21: Nuclear  
Chemistry 21.1: The  
Nature of Nuclear  
Reactions

Nucleons: - the particles that make up a nucleus of an atom (protons,  ${}^1_1\text{p}^+$  or  ${}^1_1\text{H}$ ) and neutrons,  ${}^1_0\text{n}$ ).

Isotopes: - atoms that have different mass number but the same atomic number or

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number of protons.

## Chemistry Section 1 **Chapter 21 Nuclear Chemistry Review Answers**

Nuclear Chemistry

Nuclear

Transformations •

Rutherford in 1919

performed the first

nuclear

transformation. • The

transmutations are

sometimes

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represented by listing  
in order, the target  
nucleus, the  
bombarding particle,  
the ejecting particle  
and the product  
nucleus. • The above  
equation becomes:  ${}_{14}^{27}\text{N} + {}_2^4\text{He} \rightarrow {}_{16}^{31}\text{S} + {}_0^1\text{n}$

## **Chapter 21 Nuclear Chemistry**

Nuclear chemistry is

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the study of reactions that involve changes in nuclear structure.

The chapter on atoms, molecules, and ions introduced the basic idea of nuclear structure, that the nucleus of an atom is composed of protons and, with the exception of  ${}^1_1\text{H}$ , neutrons. Recall that the number of

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protons in the nucleus is called the atomic number ( $Z$ ) of the element, and the sum of the number of protons and the number of neutrons is the mass number ( $A$ ).

### **21.1 Nuclear Structure and Stability – Chemistry**

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#### **REVIEW Nuclear**

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Chemistry SECTION

4 SHORT ANSWER

Answer the following  
questions in the

space provided. 1.

Match each of the  
following statements  
with the process(es)

to which they apply,

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using one of the  
choices below: (1)  
fission only (3) both  
fission and fusion

## **Chapter 21 Nuclear Chemistry Section 3 Answers**

Chapter 21 Nuclear  
Chemistry. Section  
21.2 Nuclear  
Reactions and  
Energy. Objectives:  
Compare and



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Contrast Nuclear fission and Fusion, Demonstrate Equations that Represent the Changes that Occur During Radioactive Decay, Trace the Operation and Structure of a Nuclear Reactor. The Power of the Nucleus. Nuclear reactions involve enormous

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energy changes.

Chemistry

### Chapter 21 Nuclear

### Chemistry -

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By the end of this section, you will be able to: Identify common particles and energies involved in nuclear reactions  
Write and balance nuclear equations  
Changes of nuclei

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that result in changes in their atomic numbers, mass numbers, or energy states are nuclear reactions.

### **21.2 Nuclear Equations – Chemistry**

Chapter 21 Nuclear Chemistry. STUDY. PLAY. nuclear reaction. a change in

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the composition or  
makeup of the  
nucleus. nuclear  
decay. the breakdown  
of nuclei by giving off  
particles or  
electromagnetic  
radiation, natural.  
fission. splitting of an  
atom into smaller  
atoms; requires lots of  
energy, not at high  
temp, chain reaction.

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Nuclear reactions

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Nuclear change one type of nucleus into another; chemical changes rearrange atoms. Nuclear reactions involve much larger energies than chemical reactions and have measureable mass changes. 9. (a), (b), (c), (d), and (e)



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## **21 - Chemistry 2e | OpenStax**

Chapter 21. Nuclear  
Chemistry. 21.5 Uses  
of Radioisotopes.

Learning Objectives.

By the end of this  
section, you will be  
able to: List common  
applications of  
radioactive isotopes;  
Radioactive isotopes  
have the same  
chemical properties

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## Chapter 21

as stable isotopes of the same element, but they emit radiation, which can be detected. If we replace one (or ...

### **21.5 Uses of Radioisotopes – Chemistry**

If the parent nuclide undergoing  $\beta^-$  decay lies below the band of stability (refer to

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Chapter 21.1 Nuclear  
Structure and  
Stability), the

daughter nuclide will  
lie closer to the band.  
Beta (?) decay is the  
emission of an  
electron from a  
nucleus. Iodine-131 is  
an example of a  
nuclide that  
undergoes ? decay:

## **21.3 Radioactive**

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692 Chapter 16

Nuclear Chemistry

16.1 The Nucleus and

Radioactivity Our

journey into the center

of the atom begins

with a brief review.

You learned in

Chapter 3 that the

protons and neutrons

in each atom are

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found in a tiny, central nucleus that measures about  $1/100,000$  the diameter of the atom itself. You also learned

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