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measurement has  
two significant

figures. The

measurement has  
four significant  
figures. Decimal  
point is

present, so

these zeros are  
significant. The  
rules are

summarized in

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the following

flowchart: qq

Nonzero digits 3

4 0 0 g of

sulfur pp

Nonzero digits 4

0 0 0. mL of

oxygen pp

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8.

Stoichiometry.

So far in your  
chemistry

course, you have

learned that

chemists count

quantities of

elements and

compounds in

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terms of moles and that they re- late moles of a substance to mass by using the molar mass. In addition, you have learned to write chemical equations so that they represent the rearrangements of atoms that

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take place

during chemical reactions, and you have learned to balance these equations.

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SAMPLE PROBLEM 1

What is the molarity of a solution that is made by diluting 50.00 mL of a 4.74 M ...

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SOLVING 1.

Determine the  
percentage

composition of

each of the

following

compounds: a.

sodium oxalate,

$\text{Na}_2\text{C}_2\text{O}_4$

ans: 34.31% Na,

17.93% C, 47.76%

O b. ethanol, C

$\text{C}_2\text{H}_5\text{OH}$  ans:

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52.13% C, 13.15%

H, 34.72% O c.

aluminum oxide,

Al<sub>2</sub>O<sub>3</sub> ans:

52.92% Al,

47.08% O d.

potassium

sulfate, K<sub>2</sub>SO<sub>4</sub>

ans: 44.87% K,

18.40% S, 36.72%

O 2.

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**SOLVING CHAPTER**

**6**

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SOLVING General  
Plan for Solving  
Molarity

Problems Mass of  
solute in g 1

Amount of solute  
in mol M moles

solute liter

solution 2

Volume of

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Guide To in L 3

Molar

concentration, M

4 Convert using

the molar mass

of the solute.

MOLARITY

Molarity is the

most common way

to express

concentration in

chemistry.

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14 . . .**

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SOLVING Neither  
pressure nor  
temperature is  
constant;  
therefore, the  
combined gas law  
must be used to  
calculate the  
volume of CO<sub>2</sub> at

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the final tem-  
perature and  
pressure.

multiply by the  
volume ratio, CO  
2 use the  
combined gas law  
to convert from  
the initial  
temperature and  
pressure to the  
final

temperature and  
pressure C 2 H 4

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Neither pressure  
nor temperature  
is constant;

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Therefore, the combined gas law must be used to calculate the volume of  $\text{CO}_2$  at the final temperature and pressure.

multiply by the volume ratio,  $\text{CO}_2$  use the combined gas law to convert from the

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empirical

formula for

compounds that

have the

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knowledgeable,

and the results

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out the

following

computations,

and express the

result in sci-

entific notation:

a.  $2.290 \times 10^7$

cm  $4.33 \times 10^3$



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ans: 5.29 !

103 cm/s b.

1.788 !  $10^5$  L "

7.111 !  $10^3$  m<sup>2</sup>

ans: 2.514 !

$10^3$  L/m<sup>2</sup> c.

5.515 !  $10^4$  L "

6.04 !  $10^3$  km

ans: 9.13 L/km

d. 3.29 !  $10^4$

km " 1.48 !  $10^2$

min ans: 2.22 !

$10^2$  km/min

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Gas Law Answers  
know that it is  
a time consuming  
job to write  
dissertations. I  
had no Chemfile  
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Answers time to  
compete my  
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but my friend  
recommended this  
website. The  
second paper I  
ordered was a  
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on history.

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**Gas Law Answers**

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PROBLEM SOLVING

3. Aluminum will  
react with

sulfuric acid in  
the following

reaction.  $2\text{Al}(s)$

$3\text{H}_2\text{SO}_4(l)$ :

$\text{Al}_2(\text{SO}_4)_3$

$(aq)$   $3\text{H}_2(g)$  a.

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How many moles  
of  $\text{H}_2\text{SO}_4$  will  
react ans: 27  
mol  $\text{H}_2\text{SO}_4$   
with 18 mol Al?

b. How many  
moles of each  
product ans: 27  
mol H will be  
produced? 2 9  
mol Al 2 ( $\text{SO}_4$ )  
3 4.

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