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~~Law Problems -~~

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Problems: Directions:
Solve the following problems. Round your answers using significant figures. 1) Calculate the mass of 15.0 L of NH_3 at 27°C and 900.0 mm Hg. 2) A volume of 26.5 mL of nitrogen gas was collected in a tube at a temperature of 17°C and a pressure of 737 ...

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Worksheet Answers

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Worksheet - Solutions

1) How many moles
of gas occupy 98 L at
a pressure of 2.8
atmospheres and a
temperature of 292
K? $n = PV = (2.8$

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atm)(98 L) = 11 moles
of gas $RT (0.0821$
 $L \cdot atm/mol \cdot K)(292 K)$
2) If 5.0 moles of O_2
and 3.0 moles of N_2
are placed in a 30.0 L
tank at a temperature
of 25 °C

Mixed Gas Laws
Worksheet Answers
Mixed Gas Laws
Worksheet - Solutions
1) How many moles

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of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K? $n = PV = (2.8 \text{ atm})(98 \text{ L}) = 11 \text{ moles}$ of gas $RT (0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})$

2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25 0

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Gas Laws Mixed

Practice Answer Key

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Answer Key

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*Gas Laws Mixed
Practice Answer Key*
The Mixed Gas Laws
Worksheet Answers
will explain the
following: how much
is in a tank, what type
of fuel is used, how
and where it is stored,
and when it is
available to be used.

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The answers that you will receive for these questions will vary from one state to another, but you will most likely receive similar answers.

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- Solutions 1) How
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K? $n = PV = (2.8$

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are

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Problems: 1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature.

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