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Atoms of a given element are identical in size, mass and other properties; atoms of different elements differ in size, mass, and other properties. 3. Atoms cannot be subdivided, created or destroyed. 4. Atoms of different elements combine in simple whole-number ratios to form chemical compounds. 5.

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3 Atoms: The Building Blocks of Matter

1: Chemistry Is a Physical Science: Section 1 Review: p.5: 2: Matter and Its Properties: Section 2 Review: p.14: 3: Elements: Section 3 Review: p.20: Chapter Review: p.22

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CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

5 The Periodic Law

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State the following measured quantities in the units indicated. a. 5.2 cm of magnesium ribbon in millimeters b. 0.049 kg of sulfur in grams c. 1.60 mL of ethanol in microliters d. 0.0025 g of vitamin A in micrograms e. 0.020 kg of tin in milligrams f. 3 kL of saline solution in liters

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4. atoms of different elements combine in simple, whole-number ratios to form compounds.

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