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Pearson Chapter 8

Covalent Bonding

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Covalent Bonding

Single Covalent

Bonds When only

one pair of

electrons is shared,

such as in a

hydrogen

molecule, it is a

single covalent

bond. The shared

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electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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configuration. In  
covalent bonding,  
atoms share  
electrons to  
achieve noble gas

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configuration. Most atoms share electrons until they have a total of 8 valence electrons (octet rule).

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covalent bonds.  
The majority

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8: Covalent

Bonding. Covalent

Bond. Molecule.

Diatomic molecule.

Molecular

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Answers

compound. A bond formed by the sharing of electrons between atoms. A neutral group of atoms joined together by covalent bonds. A molecule consisting of two atoms.

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8.2 The Nature of

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Combinations of

atoms of the

nonmetals and

metalloids in



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Answers

Groups 4A, 5A, 6A, and 7A of the periodic table are likely to form covalent bonds. □

The combined atoms usually acquire a total of eight electrons, or an

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starting the  
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answers to

admission every

daylight is

gratifying for many

people. However,

there are yet many

people who along

with don't similar

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problem. But,

bearing in mind

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others to start

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Polar Covalent

Bonds □ Though

atoms often form

compounds by

sharing electrons,

the electrons are

not always shared

equally. □ Fluorine

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Answers  
pulls harder on the electrons it shares with hydrogen than hydrogen does.

□ Therefore, the fluorine end of the molecule has more electron density than the hydrogen end.

Chapter 8  
Concepts of  
Chemical Bonding -

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Chapter 8

Chapter 8 Covalent  
Bonding 183

Section Review

Objectives □ State  
a rule that usually  
tells how many  
electrons are  
shared to form a  
covalent bond □

Describe how  
electron dot  
formulas are used  
□ Predict when two

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atoms are likely to  
be joined by a  
double or a triple  
covalent bond □

Distinguish

between a single  
covalent bond and  
other covalent  
bonds

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