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9.3 Polarity

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In polar
covalent bonds,

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the electrons are shared unequally, as one atom exerts a stronger force of attraction on the electrons than the other. The ability of an atom to attract a pair of electrons in a chemical bond is called its el

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Electronegativity. The difference in electronegativity between two atoms determines how polar a bond will be.

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for oxygen, the electronegativity difference is 1.4. We would expect a very polar bond, but not so polar that the O-H bond is considered

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and Bond Types E
lectronegativity

(EN) : The
ability for an
atom to attract
or electrons
within a bond.

Differences in
EN (Δ EN)

determine the
type of bond
that is formed
between 2 atoms.

5 Important Bond-

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form an ionic

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oxygen, the elec
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difference is
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ativity is the

tendency of a

bonded atom to

attract

electrons to

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(Delta EN)

between bonded atoms can indicate whether the bond is nonpolar, polar covalent, or ionic.

Generally, the farther apart two elements are on the ...

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more likely to

steal an

electron from
another atom.

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across period/row
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Anything greater
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anything less
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covalent. This
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What is the electronegativity of CO₂? | Study.com

The nearer the difference in electronegativity between atoms comes to zero, the purer the covalent bond becomes and the less polarity it has. Carbon, with an electron

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2.5, forms both low- and high-polarity covalent bonds. The electronegativity values of elements commonly found in organic molecules are given in Table .

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and *Electronegativity*

Electronegativity is a measure of the tendency of an atom to attract a bonding pair of electrons. The Pauling scale is the most commonly used. Fluorine (the most

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electronegative
element) is
given a value of
4.0, and values
range down to
caesium and
francium which
are the least
electronegative
at 0.7.

*electronegativity - polar bonds
in organic*

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compounds

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The electron pairs shared between two atoms are not necessarily shared equally.

Extreme

examples: 1. In Cl_2 the shared electron pairs is shared

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equally. 2. In NaCl, the 3s electron is stripped from the Na atom and is incorporated into the electronic structure of the Cl atom - and the compound is most accurately described as consisting of

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