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Antimatter Explained

Practice Problem:

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Decay Equations #34

Radioactive Decay

26 Nuclear

Equations

Writing Nuclear

Equations for Alpha

Decay Alpha Particles,

Beta Particles, Gamma

Rays, Positrons,

Electrons, Protons, and

Neutrons

Balancing nuclear

equations Alpha Decay

Writing Nuclear

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Equations Answer Key

Writing Nuclear
Equations KEY Write
nuclear equations that
describe the following
processes. 1.

Uranium-235 undergoes
an alpha decay to
produce thorium-231. 2.

Lanthanum -144
becomes cerium-144
when it undergoes a
beta decay. 3.

Neptunium-233 is

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formed when
americium-237
undergoes a nuclear
decay process. 4.

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Nuclear Equation
Practice Answer Key
Write the nuclear
equation for this
reaction and identify the
two other particles. n1n

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0 97 40 137 52 1 0 235

${}_{92}\text{U}^{235} + \text{?} \rightarrow \text{Te} + \text{Zr} + 2 \text{The}$
two other particles
formed are neutrons.

Nuclear Equation

Practice Answer Key

Nuclear Equations 1.

Bombardment of
aluminum-27 by alpha
particles produces
phosphorous-30 and one
other particle. Write the
nuclear equation for this

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reaction and identify the
other particle. ${}_{92}^{238}\text{U} \rightarrow {}_{90}^{234}\text{Th} + {}_{2}^{4}\text{He}$

The other particle
formed is a neutron n .
Plutonium-239 can be
produced by
bombarding
uranium-238 with alpha
particles.

KEY - Nuclear Equations

20 nuclear equations

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that describe the following processes. 1. Uranium-235 undergoes an alpha decay to produce thorium-231. 2. Lanthanum -144 becomes cerium-144 when it undergoes a beta decay. 3. Neptunium-233 is formed when americium-237 undergoes

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**Nuclear Equation
Practice Answer Key -
Orris**

**NUCLEAR
EQUATIONS
WORKSHEET**

ANSWERS 1. Write a nuclear equation for the alpha decay of $^{231}_{91}\text{Pa}$

$^{231}_{91}\text{Pa} \rightarrow ^{4}_{2}\text{He} + ^{227}_{89}\text{Ac}$

2. Write a nuclear equation for the beta decay of $^{223}_{87}\text{Fr}$

$^{223}_{87}\text{Fr} \rightarrow ^{0}_{-1}\text{e} + ^{223}_{88}\text{Ra}$

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88.3. Write a nuclear equation for the alpha and beta decay of $^{149}_{62}\text{Sm}$.
 $^{149}_{62}\text{Sm} \rightarrow ^4_2\text{He} + ^{145}_{60}\text{Nd} + ^0_{-1}\text{e}$

NUCLEAR EQUATIONS WORKSHEET ANSWERS

Write a balanced nuclear equation for a natural transmutation.

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Prerequisites atomic
symbols including mass
number, atomic number,
and charge Information

4He = a helium
nucleus, also known as
an alpha particle = an
electron, also known as
a beta particle when
emitted by a nucleus

Model The following
are two nuclear reaction
equations: $4\text{He} + 216$

$\text{At } 85 \text{ } 87 \text{ } 16 \text{ IV } \dots + +$

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Equations

Ms. Demonte's Answer Key Chemistry Classes - Home

Identify common particles and energies involved in nuclear reactions. Write and balance nuclear equations. Changes of nuclei that result in changes in their atomic numbers, mass numbers,

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Nuclear energy states are nuclear reactions. To describe a nuclear reaction, we use an equation that identifies the nuclides involved in the reaction, their mass numbers and atomic numbers, and the other particles involved in the reaction.

21.2 Nuclear Equations – Chemistry

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12. Write a balanced nuclear equation for each decay process indicated.

a. The isotope Th-234 decays by an alpha emission.

b. The isotope Fe-59 decays by a beta emission.

c. The isotope Tc-99 decays by a gamma emission.

d. The isotope C-11 decays by a electron capture.

Balance these equations:

Note ! $2\text{ }^4_2\text{He}$ is the only

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stable isotope of helium

... Equations

Answer Key

Balancing Nuclear Equations

${}_{92}^{238}\text{U}$. That 92 is the atomic number, which is the number of protons.

That 238 is the mass number, which is the sum of the protons and the neutrons. Lastly, remember that you have to do a subtraction to get

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the number of neutrons:
 $14 - 6 = 8$. $238 - 92 =$
146.

Answer Key

ChemTeam: Writing Alpha and Beta Equations

Write the complete
nuclear equation. beta
particle (e^-) + N-14.
beta particle (e^-) +
Y-90. beta particle (e^-)
+ Ca-40. beta particle
(e^-) + O-13. The

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following all undergo
electron capture. Write
the complete nuclear
equation. electron (e-) ?
Pd-106. + electron (e-) ?
In-116.

Nuclear decay worksheet - CTE Online

Writing nuclear
equations for alpha,
beta, and gamma decay.
... So for representing an

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alpha particle in our nuclear equation, since an alpha particle has the same composition as a helium nucleus, we put an He in here, and it has two positive charges, so we put a two down here, and then a total of four nucleons, so we put a four here. ...

**Writing nuclear
equations for alpha,**

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beta, and gamma decay...

After completing this I
will again take

volunteers to write the
problems from this
section on the board as
seen on the answer key.
Students generally do
very well on this
activity and learn the
basic format for writing
a nuclear equation. This
can be seen in the EL

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Nuclear Decay student
work this provided. This
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foundation ...

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Answer Key Chemistry

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Nuclear Equations KEY

Write nuclear equations
that describe the

following processes. 1.

Uranium-235 undergoes
an alpha decay to

produce thorium-231. 2.

Lanthanum -144

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becomes cerium-144
when it undergoes a
beta decay. 3.

Neptunium-233 is
formed when Page 3/23

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